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table, he was able to predict the ...

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## Answers

Section 5.1 - Organizing the Elements. A periodic table is an arrangement of elements in columns, based on a set of properties that repeat from row to row. Dmitri . Mendeleev. is credited with creating the first useful . periodic table. Mendeleev. arranged the elements into rows in order of increasing . mass

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Concepts: Terms in this set (46) 1. Why did  
Mendeleev not include helium in his periodic  
table of 1871. 1. It was not discovered yet

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## Answers

(discovered in 1895) 3. Why did Mendeleev not include radium in his periodic table of ...

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CHAPTER 5 REVIEW The Periodic Law SECTION 1

SHORT ANSWER Answer the following questions

in the space provided. 1. c In the modern

periodic table, elements are ordered (a)

according to decreasing atomic mass. (b)

according to Mendeleev's original design. (c)

according to increasing atomic number. (d)

based on when they were discovered.

### **5 The Periodic Law**

The left side of the periodic table is

composed of metals, which have few valence

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## Answers

electrons. Moving to the right, there are more electrons and their ... ANSWERS To TOPIC 5 Review Questions 12. 21. 24. 27. 30. 33. 36. 39. 42. 45. 48. 51. 54. 57. 60. 63. 66. 69. 72. 4 3 3 2 4 4 4 1 1 1 4 2 3 3 1 1 1 4 3 1. 10. 13. 16. 19. 22. 25.

### **Answers to Topic 5 Periodic Table from Review Book**

- The elements of Group 2 of the periodic table are called the alkaline-earth metals. • beryllium, magnesium, calcium, strontium, barium, and radium

Section 2 Electron Configuration and the Periodic Table Chapter

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## Answers

5 • Group 2 metals are less reactive than the alkali metals, but are still too reactive to be found in nature in pure form.

### **The Periodic Table.ppt - Chapter 5 Section 1 History of ...**

The Periodic Table of the Elements is useful for revealing patterns and trends in the elements. Which statement accurately describes a pattern in the size of atomic radii in the Periodic Table of the Elements?

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Chapter 5 5 Refined Arrangement • H. G. J. Moseley discovered that the nuclear charge increased by one for each element on the periodic table. • He concluded that if the elements are arranged by increasing nuclear charge rather than atomic mass, the trends on the periodic table are better explained.

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## Answers

periodic table, classify each of the following elements as a metal or a nonmetal, and then further classify each as a main-group (representative) element, transition metal, or inner transition metal: (a) uranium (b) bromine (c) strontium (d) neon (e) gold (f) americium

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