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Chapter 6 – Chemical Bonds

Writing Formulas for Ionic Compounds What is the ratio of the ions in magnesium iodide? What is the formula for magnesium iodide? 1. Read and Understand

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This textbook introduces the molecular and quantum chemistry needed to understand the physical properties of molecules and their chemical bonds. It follows the authors' earlier textbook "The Physics of Atoms and Quanta" and presents both experimental and theoretical fundamentals for students in physics and physical and theoretical chemistry. The new edition treats new developments in areas such as high-resolution two-photon spectroscopy, ultrashort pulse spectroscopy, photoelectron spectroscopy, optical investigation of single molecules in condensed phase, electroluminescence, and light-emitting diodes.

Winner of the PROSE Award for Chemistry & Physics 2010 Acknowledging the very best in professional and scholarly publishing, the annual PROSE Awards recognise publishers' and authors' commitment to pioneering works of research and for contributing to the conception, production, and design of landmark works in their fields. Judged by peer publishers, librarians, and medical professionals, Wiley are pleased to congratulate Professor Ian Fleming, winner of the PROSE Award in Chemistry and Physics for Molecular Orbitals and Organic Chemical Reactions. Molecular orbital theory is used by chemists to describe the arrangement of electrons in chemical structures. It is also a theory capable of giving some insight into the forces involved in the making and breaking of chemical bonds—the chemical reactions that are often the focus of an organic chemist's interest. Organic chemists with a serious interest in understanding and explaining their work usually express their ideas in molecular orbital terms, so much so that it is now an essential component of every organic chemist's skills to have some acquaintance with molecular orbital theory. Molecular Orbitals and Organic Chemical Reactions is both a simplified account of molecular orbital theory and a review of its applications in organic chemistry; it provides a basic introduction to the subject and a wealth of illustrative examples. In this book molecular orbital theory is presented in a much simplified, and entirely non-mathematical language, accessible to every organic chemist, whether student or research worker, whether mathematically competent or not. Topics covered include: Molecular Orbital Theory Molecular Orbitals and the Structures of Organic Molecules Chemical Reactions — How Far and How Fast Ionic Reactions — Reactivity Ionic Reactions — Stereochemistry Pericyclic Reactions Radical Reactions Photochemical Reactions This expanded Reference Edition of Molecular Orbitals and Organic Chemical Reactions takes the content and the same non-mathematical approach of the Student Edition, and adds extensive extra subject coverage, detail and over 1500 references. The additional material adds a deeper understanding of the models used, and includes a broader range of applications and case studies. Providing a complete in-depth reference for a more advanced audience, this edition will find a place on the bookshelves of researchers and advanced students of organic, physical organic and computational chemistry. The student edition of Molecular Orbitals and Organic Chemical Reactions presents molecular orbital theory in a simplified form, and offers an invaluable first textbook on this important subject for students of organic, physical organic and computational chemistry. Further information can be viewed here. "These books are the result of years of work, which began as an attempt to write a second edition of my 1976 book Frontier Orbitals and Organic Chemical Reactions. I wanted to give a rather more thorough introduction to molecular orbitals, while maintaining my focus on the organic chemist who did not want a mathematical account, but still wanted to understand organic chemistry at a physical level. I'm delighted to win this prize, and hope a new generation of chemists will benefit from these books." —Professor Ian Fleming

In his highly anticipated sequel to The Elements, Theodore Gray demonstrates how the elements of the periodic table combine to form the molecules that make up our world. Everything physical is made up of the elements and the infinite variety of molecules they form when they combine with each other. In Molecules, Theodore Gray takes the next step in the grand story that began with the periodic table in his best-selling book, The Elements: A Visual Exploration of Every Known Atom in the Universe. Here, he explores through fascinating stories and trademark stunning photography the most interesting, essential, useful, and beautiful of the millions of chemical structures that make up every material in the world. Gray begins with an explanation of how atoms bond to form molecules and compounds, as well as the difference between organic and inorganic chemistry. He then goes on to explore the vast array of materials molecules can create, including: soaps and solvents; goops and oils; rocks and ores; ropes and fibers; painkillers and dangerous drugs; sweeteners; perfumes and stink bombs; colors and pigments; and controversial compounds including asbestos, CFCs, and thimerosal. Big, gorgeous photographs, as well as diagrams of the compounds and their chemical bonds, rendered with never before seen beauty, fill the pages and capture molecules in their various states. As he did in The Elements, Gray shows us molecules as we've never seen them before. It's the perfect book for his loyal fans who've been eager for more and for anyone fascinated with the mysteries of the material world.

Physics is a complex and daunting topic, but it is also deeply satisfying--even thrilling. And it is absolutely one you can understand. Absolutely Small develops your intuition for the very nature of things at their most basic and intriguing levels by demystifying the world of quantum science. Just as we can understand the concept of gravity without solving a single equation, author Michael D. Fayer, professor of chemistry at Stanford University, uses examples from the everyday world to help you understand quantum science like never before. Exploring a range of scientific concepts--from particles of light, to probability, to states of matter, to what makes greenhouse gases bad--in considerable depth, he provides readers the answers to questions like: What makes blueberries blue and strawberries red? Does sound really travel in waves? and Why does light behave so differently from any other phenomenon in the universe? Challenging without being intimidating and accessible but not condescending, Absolutely Small liberates physics from its mathematical underpinnings so anyone with curiosity and imagination can explore its beauty.

One of The New York Times Book Review 's " 10 Best Books of 2021 " Shortlisted for the 2021 International Booker Prize A fictional examination of the lives of real-life scientists and thinkers whose discoveries resulted in moral consequences beyond their imagining. When We Cease to Understand the World is a book about the complicated links between scientific and mathematical discovery, madness, and destruction. Fritz Haber, Alexander Grothendieck, Werner Heisenberg, Erwin Schr ö dinger—these are some of luminaries into whose troubled lives Benjam i n Labatut thrusts the reader, showing us how they grappled with the most profound questions of existence. They have strokes of unparalleled genius, alienate friends and lovers, descend into isolation and insanity. Some of their discoveries reshape human life for the better; others pave the way to chaos and unimaginable suffering. The lines are never clear. At a breakneck pace and with a wealth of disturbing detail, Labatut uses the imaginative resources of fiction to tell the stories of the scientists and mathematicians who expanded our notions of the possible.

"Understanding Information "illustrates the basic principles of information science, to provide a general introduction to the subject, through a series of selected and interesting examples. It touches on a variety of issues, including Intranets and knowledge management. All those who are involved in the turbulent changes in the information field will find a picture of how information and its concepts operate in contemporary society.

Discusses the Structure and Properties of Materials and How These Materials Are Used in Diverse Applications Building on undergraduate students ' backgrounds in mathematics, science, and engineering, Introduction to the Physics and Chemistry of Materials provides the foundation needed for more advanced work in materials science. Ideal for a two-semester course, the text focuses on chemical bonding, crystal structure, mechanical properties, phase transformations, and materials processing for the first semester. The material for the second semester covers thermal, electronic, photonic, optical, and magnetic properties of materials. Requiring no prior experience in modern physics and quantum mechanics, the book introduces quantum concepts and wave mechanics through a simple derivation of the Schr ö dinger equation, the electron-in-a-box problem, and the wave functions of the hydrogen atom. The author also presents a historical perspective on the development of the materials science field. He discusses the Bose – Einstein, Maxwell – Boltzmann, Planck, and Fermi – Dirac distribution functions, before moving on to the various properties and applications of materials. With detailed derivations of important equations, this applications-oriented text examines the structure and properties of materials, such as heavy metal glasses and superconductors. It also explores recent developments in organics electronics, polymer light-emitting diodes, superconductivity, and more.

EDITIONS: This book is available in paperback in 5.5" x 8.5" (portable size), 8.5" x 11" (large size), and as an eBook. This 5.5" x 8.5" edition is the most portable, while the details of the figures - including the periodic tables - are most clear in the large size and large print edition. However, the paperback editions are in black-and-white, whereas the eBooks are in color. OVERVIEW: This book focuses on fundamental chemistry concepts, such as understanding the periodic table of the elements and how chemical bonds are formed. No prior knowledge of chemistry is assumed. The mathematical component involves only basic arithmetic. The content is much more conceptual than mathematical. AUDIENCE: It is geared toward helping anyone – student or not – to understand the main ideas of chemistry. Both students and non-students may find it helpful to be able to focus on understanding the main concepts without the constant emphasis on computations that is generally found in chemistry lectures and textbooks. CONTENTS: (1) Understanding the organization of the periodic table, including trends and patterns. (2) Understanding ionic and covalent bonds and how they are formed, including the structure of valence electrons. (3) A set of rules to follow to speak the language of chemistry fluently: How to name compounds when different types of compounds follow different naming schemes. (4) Understanding chemical reactions, including how to balance them and a survey of important reactions. (5) Understanding the three phases of matter: properties of matter, amorphous and crystalline solids, ideal gases, liquids, solutions, and acids/bases. (6) Understanding atomic and nuclear structure and how it relates to chemistry. (7) VERBAL ReActiONS: A brief fun diversion from science for the verbal side of the brain, using symbols from chemistry's periodic table to make word puzzles. ANSWERS: Every chapter includes self-check exercises to offer practice and help the reader check his or her understanding. 100% of the exercises have answers at the back of the book. COPYRIGHT: Teachers who purchase one copy of this book or borrow one copy of this book from a library may reproduce selected pages for the purpose of teaching chemistry concepts to their own students.

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